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[NAMING IONIC COMPOUNDS WORKSHEET 2](#)

Naming Ionic Compounds Worksheet 2 W 308 Name the following ionic compounds: 1) $\text{Mo}(\text{SO}_4)_3$ 2) $\text{Ti}(\text{NO}_3)_3$ 3) NaH 4) $\text{BeSO}_3 \cdot 7\text{H}_2\text{O}$ 5) $\text{Fe}_3(\text{PO}_4)_2$ 6) $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ 7) $\text{V}(\text{H}_2\text{PO}_4)_4$ 8) $\text{Mn}_3(\text{PO}_4)_7$ 9) $\text{Fe}(\text{CN})_3$ 10) BaSO_4 Write the formulas for the following ionic compounds: 11) vanadium (V) arsenide 12) cobalt (III) nitrite 13) yttrium perchlorate 14) silver phosphide 15) tin (IV) selenide 16) cadmium sulfite 17) nickel (III)...

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AP Chemistry Chapter 12 Answers Zumdahl 12.49 For elementary reactions, the rate law can be written using the coefficients in the balanced equation to determine orders. a. $\text{Rate} = k[\text{CH}_3\text{NC}]$ b. $\text{Rate} = k[\text{O}_3][\text{NO}]$ c. $\text{Rate} = k[\text{O}_3]$ d. $\text{Rate} = k[\text{O}_3][\text{O}]$ 12.50 The observed rate law for this reaction is: $\text{Rate} = k[\text{NO}]^2[\text{H}_2]$. For a mechanism to be plausible, the sum of all the steps must give the overall balanced equation...

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[MICROSOFT WORD - CH. 14 SECTIONS 14.1-14.2 TEXTBOOK HOMEWORK ANSWERS](#)

AP Chemistry Chapter 14 Answers Zumdahl 14.27 a. $\text{HClO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. Only the forward reaction is indicated since HClO_4 is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is: $\text{HClO}_4(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. This reaction is also called the K_a re...

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