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Honour Chemistry: Unit 4 Practice Test: Thermochemistry and Nuclear ChemistryPart A: Multiple Choice and Numerical Response (1 mark each for Multiple Choice) (2 marks each for Numerical Response) $E_{\text{bind}} = \sum mc^2$ $c = \text{speed of light}$ (3.00 10⁸ m s⁻¹) 1 eV = 1.69 10⁻¹⁹ J (1 kg = 6.022 10²⁶ amu , or 1 g = 6.022 10²³ amu = 1 mole amu) 1 MeV = 1.69 10⁻¹⁶ J (1 Neutron = 1.008665 amu) (1 Proton = 1.007825 amu)

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[UNIT 2 \(PART 1\) PRACTICE TEST](#)

Honour Chemistry Practice Test: Unit 2 (Part 1): Matter as Solutions and Gases 1 atm = 760 mm Hg = 760 torr = 101.325 kPa $T_K = T_C + 273.15$ STP = 22.4 L mol @ 0.00C & 101.325 kPa (1 atm) SATP = 24.8 L mol @ 25.0C & 100. kPa L atm L kPa m m nsolute $R = 0.0821$ or 8.314 $D = n = C = K$ mol K mol V M Vsolution $P_1V_1 = P_2V_2$ $r_1M_2 = n_1P_1 = C_1V_1 = C_2V_2$ $PV = nRT$ $n_1T_1 = n_2T_2$ $r_2M_1 = n_{\text{Total}}P_{\text{Total}}$ $P_{\text{Total}} = P_1 + \dots$

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[PRACTICE TEST 7](#)

Practice Test 7, solutions belowProblem 11. (p & q) A2. (p v q) A3. Show: rProblem 21. (p ? q) ? (r ? s) A2. p v q A3 (q & p) A4. Show: (r v s)Problem 31. (p v q) v (r v s) A2. r A3. (t ? s) A4. Show p ?q Problem 11. (p & q) A2. (p v q) A3. Show: r4. p & q &, 25. p & q , 46. p & q , 57. r !, 1,6Problem 21. (p ? q) ? (r ? s) A2. p v q A3 (q & p) A4. Show: (r v s)5. (r v s) AIP6. p ? q ?v, 27. q v ...





















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